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In a tetrahedral molecular geometry, a central atom is located at the center with four substituents that are located at the corners of a tetrahedron. The bond angles are $\cos^{-1}(\frac{1}{\sqrt{3}}) = 109.4712206...^\circ \hat{\approx} 109.5^\circ$ when all four substituents are the same, as in methane (CH₄) as well as its heavier analogues. Methane and other perfectly symmetrical tetrahedral molecules belong to point group ...

Tetrahedral molecular geometry - Wikipedia

In chemistry, the term "bent" can be applied to certain molecules to describe their molecular geometry. Certain atoms, such as oxygen, will almost always set their two (or more) covalent bonds in non-collinear directions due to their electron configuration. Water (H₂O) is an example of a bent molecule, as well as its analogues. The bond angle between the two hydrogen atoms is approximately 104 ...

Bent molecular geometry - Wikipedia

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Chapter 9

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